## UNIVERSITY OF MYSORE Postgraduate Entrance Examination October-2021



(Read carefully the instructions given in the Question Booklet)


MAXIMUM MARKS : 50
(Including time for filling O.M.R. Answer sheet)

## INSTRUCTIONS TO THE CANDIDATES

1. The sealed question paper booklet containing 50 questions enclosed with O.M.R. Answer Sheet is given to you.
2. Verify whether the given question booklet is of the same subject which you have opted for examination.
3. Open the question paper seal carefully and take out the enclosed O.M.R. Answer Sheet outside the question booklet and fill up the general information in the O.M.R. Answer sheet. If you fail to fill up the details in the form as instructed, you will be personally responsible for consequences arising during evaluating your Answer Sheet.
4. During the examination:
a) Read each question carefully.
b) Determine the Most appropriate/correct answer from the four available choices given under each question.
c) Completely darken the relevant circle against the Question in the O.M.R. Answer Sheet. For example, in the question paper if "C" is correct answer for Question No.8, then darken against SI. No. 8 of O.M.R. Answer Sheet using Blue/Black Ball Point Pen as follows:

Question No. 8. (A) (B) (D) (Only example) (Use Ball Pen only)
5. Rough work should be done only on the blank space provided in the Question Booklet. Rough work should not be done on the O.M.R. Answer Sheet.
6. If more than one circle is darkened for a given question, such answer is treated as wrong and no mark will be given. See the example in the O.M.R. Sheet.
7. The candidate and the Room Supervisor should sign in the O.M.R. Sheet at the specified place.
8. Candidate should return the original O.M.R. Answer Sheet and the university copy to the Room Supervisor after the examination.
9. Candidate can carry the question booklet and the candidate copy of the O.M.R. Sheet.
10. The calculator, pager and mobile phone are not allowed inside the examination hall.
11. If a candidate is found committing malpractice, such a candidate shall not be considered for admission to the course and action against such candidate will be taken as per rules.
12. Candidates have to get qualified in the respective entrance examination by securing a minimum of 8 marks in case of SC/ST/Cat-I Candidates, 9 marks in case of OBC Candidates and 10 marks in case of other Candidates out of 50 marks.

## INSTRUCTIONS TO FILL UP THE O.M.R. SHEET

1. There is only one most appropriate/correct answer for each question.
2. For each question, only one circle must be darkened with BLUE or BLACK ball point pen only. Do not try to alter it.
3. Circle should be darkened completely so that the alphabet inside it is not visible.
4. Do not make any unnecessary marks on O.M.R. Sheet.
5. Mention the number of questions answered in the appropriate space provided in the O.M.R. sheet otherwise O.M.R. sheet will not be subjected for evaluation.

1) Molecular formula of fructose is
(A) $\mathrm{C}_{12} \mathrm{H}_{22} \mathrm{O}_{11}$
(B) $\mathrm{C}_{18} \mathrm{H}_{32} \mathrm{O}_{16}$
(C) $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(D) None of these
2) Starch is an example of
(A) Monosaccharides
(B) Oligosaccharides
(C) Polysaccharides
(D) Lipids
3) Glucose is a monosaccharide and is a
(A) Hexose
(B) Pentose
(C) Furanose
(D) Sucrose
4) Which Biomolecules simply refers as "Staff of life" in the given $\qquad$
(A) Protein
(B) Lipids
(C) Carbohydrate
(D) Vitamins
5) In carbohydrates which are the main functional groups are present?
(A) Alcohol \& Carboxyl groups
(B) Aldehyde \& Ketone groups
(C) Hydroxyl groups \& Hydrogen groups
(D) Carboxyl groups \& Others
6) Which of the following is not a disaccharide?
(A) Sucrose
(B) Maltose
(C) Lactose
(D) Galactose
7) Which sugars are present in Sucrose?
(A) Fructose and glucose
(B) Glucose and glucose
(C) Glucose and galatose
(D) Fructose and galatose
8) Which one is correctly matched?
(A) Acids - pH range above 7
(B) Acids - pH range below 7
(C) Acids - pH range 7 (neutral)
(D) Acids - pH range 8-9
9) Which one will change from red litmus to blue?
(A) NaCl
(B) HCl
(C) KOH
(D) LiOH
10) Which of the following amino acids has a net negative charge at physiologic $\mathrm{pH}(\sim 7.4) ?$
(A) Glutamic Acid
(B) Lysine
(C) Histidine
(D) None of these
11) Amino acids are joined together via which bond?
(A) Dipole-Dipole
(B) Amide
(C) Hydrogen
(D) Ionic
12) The sulphuric acid turns blue litmus paper into
(A) Red
(B) Green
(C) Yellow
(D) Dark blue
13) Which one of the following metal is used thermometers?
(A) Copper
(B) Mercury
(C) Aluminium
(D) Iron
14) Heat conduction is the property of
(A) Non-metal
(B) Metal
(C) Metalloids
(D) All of these
15) What is Reynolds number?
(A) Ratio of inertial force to viscous force
(B) Ratio of weight of fluids viscous force
(C) Ratio of velocity gradient to viscous force
(D) Reciprocal of viscous force
16) Which of these fluids has the highest viscosity?
(A) Water
(B) Honey
(C) Blood
(D) Air
17) Mathematical expression that describes Boyle's law is
(A) $\mathrm{PV}=$ constant
(B) $\mathrm{V}^{*}$ constant $=\mathrm{P}$
(C) $\mathrm{P} *$ constant $=\mathrm{V}$
(D) $\mathrm{V} / \mathrm{P}=$ constant
18) In solutions particles are
(A) Invisible
(B) Visible by naked eye
(C) Visible by ordinary microscope
(D) Visible by electron microscope
19) When tiny particles of a substance are dispersed through medium then mixture is called
(A) alloys
(B) amalgams
(C) suspension
(D) colloid
20) Colloids can
(A) Absorb heat
(B) Not scatter light
(C) Scatter light
(D) Evolveheat
21) Concentration of a dilute solution
(A) is high
(B) is in comparable
(C) can be both A and B
(D) is low
22) Absolute zero on Kelvin scale is equal to
(A) 373 K
(B) 273 K
(C) 0 K
(D) None of the above
23) Which one of these thermometers is portable as well as simple to use?
(A) Constant-volume gas thermometer
(B) Resistance thermometer
(C) Thermocouple
(D) Mercury-in-glass thermometer
24) Convert a temperature measurement of 250 deg C into Kelvin.
(A) 523.2 K
(B) -209.7 K
(C) 709.7 K
(D) -23.2 K
25) Specific heat capacity of a substance is equal to
(A) Mass of the substance $\times$ heat capacity
(B) Heat capacity / mass of the substance
(C) Mass of the substance / heat capacity
(D) None of the above
26) Absolute pressure is:
(A) Gauge pressure plus atmospheric pressure
(B) Gauge pressure less atmospheric pressure
(C) Gauge pressure plus atmospheric pressure divided by two
(D) Always referenced to a point at the peak of Mt. Washington, N H
27) Every atom consists of electrons, protons and neutrons except:
(A) Helium atom
(B) Ordinary hydrogen atom
(C) Sodium atom
(D) Boron atom
28) $X$-ray beam is
(A) Electromagnetic radiation
(B) Emitted from atomic molecules
(C) Can penetrate through walls
(D) Made up of electrons
29) Maximum number of electrons in an orbit is given by
(A) $n^{2}$
(B) $2 n^{2}$
(C) $n^{2} / 2$
(D) None
30) The first scientist to publish on the periodic table was
(A) Moseley
(B) Mendeleev
(C) Fischer
(D) Dalton
31) The most electronegative element is
(A) Oxygen
(B) Sodium
(C) Iodine
(D) Fluorine
32) Which of the following substance has the least ionic character in its bonds?
(A) KCl
(B) $\mathrm{CCl}_{4}$
(C) NaCl
(D) $\mathrm{MgCl}_{2}$
33) The rate of flow of heat in a conductor depends on
(A) Temperature difference
(B) Area of cross section
(C) Length
(D) All of these
34) Green pigment present in sugarcane leaf
(A) Xanthophyll
(B) Carotene
(C) Chlorophyll
(D) Anthocyanin
35) Abundant natural source of Vitamin $C$ is
(A) Carrot
(B) Papaya
(C) Amla
(D) Banana
36) Which among the following is known as universal solvent?
(A) Methanol
(B) Hexane
(C) Water
(D) Benzene
37) In ice, the number of hydrogen bonds per water molecule is
(A) 3
(B) 4
(C) 2
(D) 1
38) The pH of pure water is
(A) 0
(B) 7
(C) 9
(D) 14
39) The oxidation state of carbon in sucrose is
(A) 4
(B) 2
(C) 1
(D) 0
40) Which one of the following substances is the main constituent of bones?
(A) Calcium phosphate
(B) Calcium chloride
(C) Calcium sulphate
(D) Calcium
41) When sodium hydroxide is dissolved in water, solution becomes
(A) Cold
(B) Hot
(C) Remains at original temperature
(D) Depends on the altitude of the place
42) Which among the following is a strong acid?
(A) $\mathrm{CH}_{3} \mathrm{COOH}$
(B) HCOOH
(C) HF
(D) $\mathrm{C}_{6} \mathrm{H}_{5} \mathrm{COOH}$
43) Ozone depletion in the stratosphere is mainly caused by :
(A) $\mathrm{SO}_{2}$
(B) NO
(C) CFCs
(D) $\mathrm{NO}_{2}$
44) The bond angle in $\mathrm{sp}^{3}$ hybrid orbitals is
(A) $120^{\circ}$
(B) $109^{\circ} 28^{\prime}$
(C) $105^{\circ}$
(D) $180^{\circ}$
45) The number of covalent bonds each carbon atom has in organic compounds is usually
(A) 4
(B) 3
(C) 2
(D) 6
46) Resonance is due to
(A) Delocalisation of sigma electrons
(B) Delocalisation of pi electrons
(C) Migration of H atoms
(D) Migration of proton
47) The density of water at $25^{\circ} \mathrm{C}$ is
(A) $0.234 \mathrm{~g} / \mathrm{mL}$
(B) $0.512 \mathrm{~g} / \mathrm{mL}$
(C) $0.998 \mathrm{~g} / \mathrm{mL}$
(D) $1.234 \mathrm{~g} / \mathrm{mL}$
48) The first organic compound synthesised in the laboratory was
(A) Urea
(B) Glucose
(C) Methane
(D) Acetic acid
49) The inverse of resistivity is called
(A) Conductivity
(B) Specific resistance
(C) Potential difference
(D) None
50) If there are $n$ asymmetric carbon atoms in a compound then the maximum number of stereoisomers that can exist is
(A) $n^{2}$
(B) $2^{n}$
(C) $2^{\mathrm{n}+1}$
(D) $2 n^{2}$

## Rough Work

## అభ్యథిรగษిగి శ్జอఒసేగఆు



 ఎంబదన్ను யరిరిలలిసిరి.



 జదాబ్దారరంగిరుత్తిర.


 లుత్తరహస్ను నిధణరిి.


 కుంబిర:




 ळలళెయల్లిన లుదాळరణ నైలణి.
 యృడ్బొలు.
 పిల్టలిద్యానిలయుద
 ஹృఁగబळుదు.




 అంచగఆన్ను யֹడియత్ర్ప్దు.

## ఓ.ఎం.ఆరా. ळలఆయన్ను కుంబలు ష్యృజసెగళు









Note : English version of the instructions is printed on the front cover of this booklet.

